

Type II Binary Ionic Compounds

transition metals 2 M : NM

- Transition metals have more than one charge
- You are told the charge by

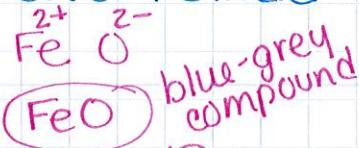
Roman numerals in the name

1	I	5	IV	9	IX
2	II	6	V	10	X
3	III	7	VI		
4	IV	8	VII		

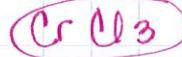
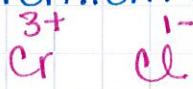
Writing formulas

- Write the symbol & charge of the metal cation (Roman numeral)
- Do the same for nonmetal anion (on PT).
- Criss-cross charges.

Ex) iron(II) oxide



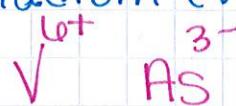
chromium(III) chloride



uranium(III) phosphide



vanadium(VI) arsenide



simplif!

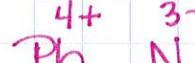
iron(III) oxide



manganese(IV) sulfide



lead(IV) nitride

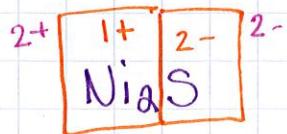


Writing Names

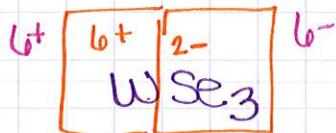
- write the name of the metal
- write the metal's original charge as a Roman numeral.
- write the nonmetal's name, change the ending to -ide



cobalt (III)
bromide



nickel (I)
sulfide



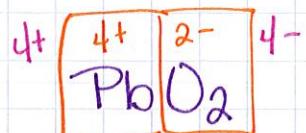
tungsten (VI)
selenide



nickel (II)
sulfide



tin (II)
fluoride



lead (IV)
oxide

