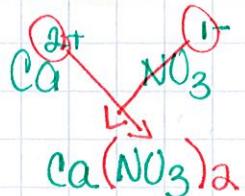


Type I & II Tertiary Ionic Nomenclature

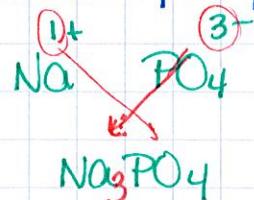
Type I Tertiary Formulas

- write the symbol & charge of metal cation
- do the same for the polyatomic ion
- criss-cross charges, use parentheses around the polyatomic ion IF you are writing a number after it.

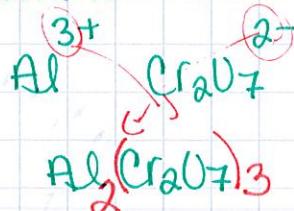
Ex.) calcium nitrate



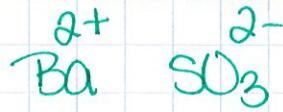
sodium phosphate



aluminum dichromate



barium sulfite



Type I Tertiary Names

- write the name of the metal cation
- write the name of the polyatomic ion.

Ex) NaOH

sodium
hydroxide

$\text{Sr}(\text{ClO}_4)_2$
strontium
perchlorate

$\text{Mg}_3(\text{PO}_4)_2$
magnesium
phosphate

NaHSO_4
sodium
hydrogen
sulfate

Type II Tertiary Formulas

has Roman numerals for transition metals (charge)

- Do the same thing you did for type I

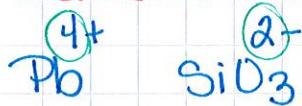
Ex) chromium (III) carbonate mercury (I) cyanide



iron (II) chromate



lead (IV) silicate

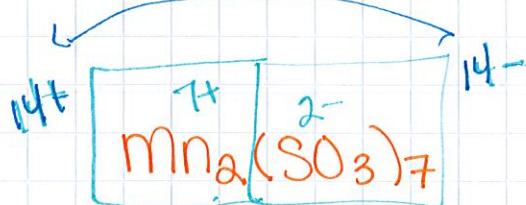


Type II Tertiary Names

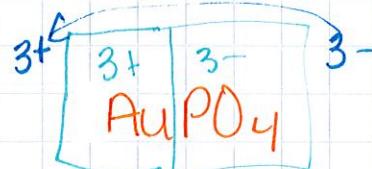
- write name of transition metal cation
- write the original charge of transition metal as a roman numeral
- write name of polyatomic ion.



zinc (II) nitrite



manganese (VII)
sulfite



gold (III)
phosphate



osmium (VI)
silicate