

# Ionic Nomenclature

## Namis & Formulas

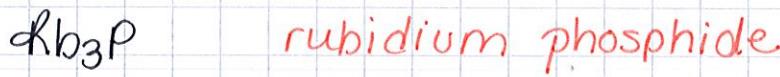
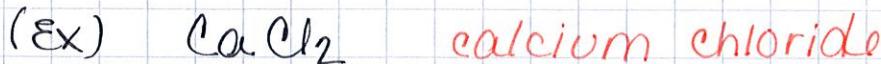
IUPAC nomenclature

International Union of Pure & Applied Chemistry

### 1. Type I Binary Ionic Nomenclature no transition & elements metals

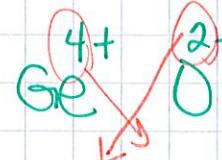
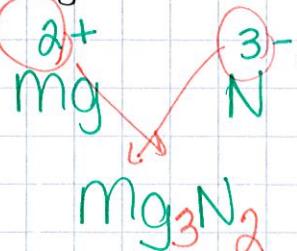
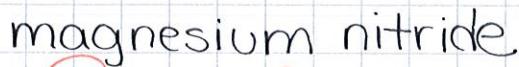
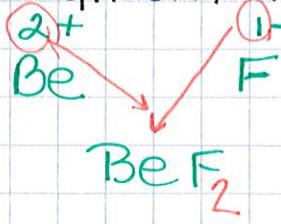
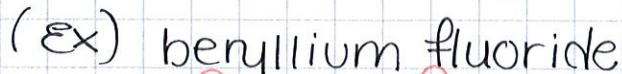
#### (A) Names

- write the name of the metal (1<sup>st</sup> element)
- write the name of the nonmetal (2<sup>nd</sup> element)  
& change its ending to -ide



#### (B) Formulas

- write the symbol & charges for the metal
- do the same for - the nonmetal
- criss-cross the charges, drop the +/ - sign & simplify if possible

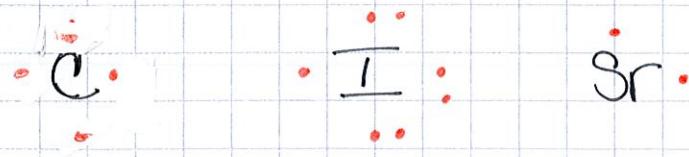


simplify!  
 $\downarrow$

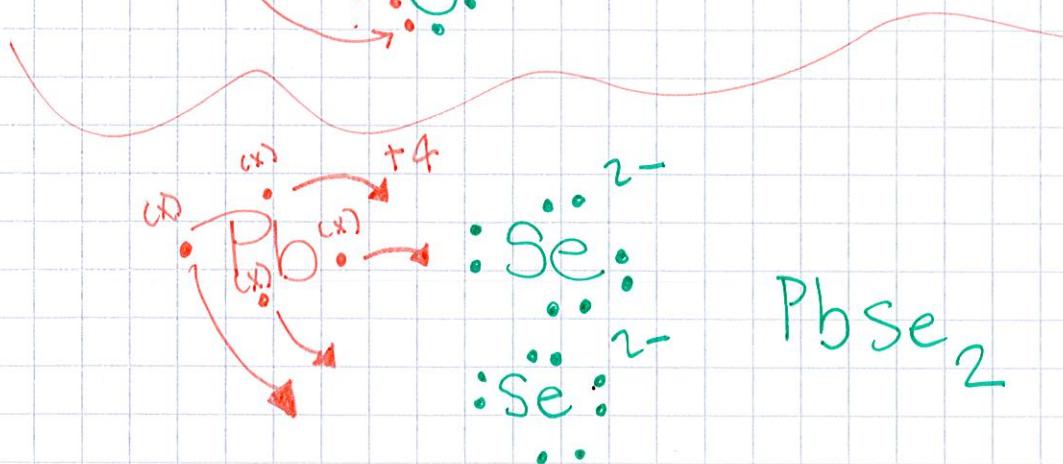
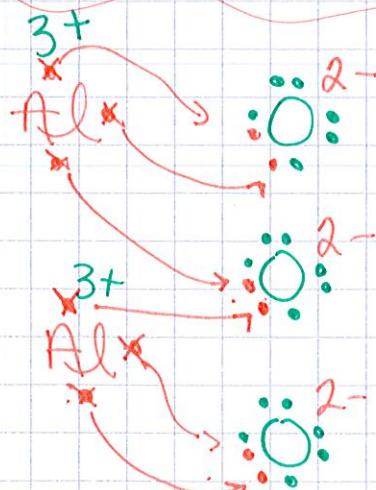
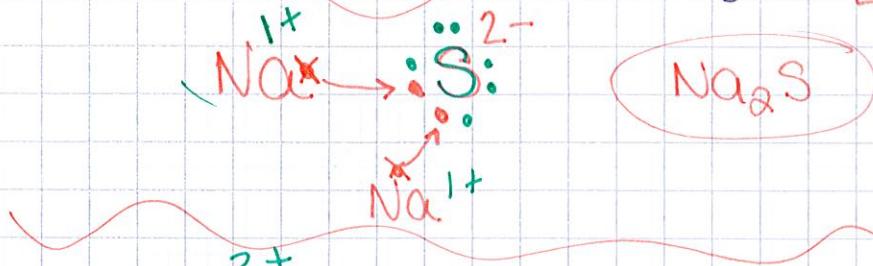
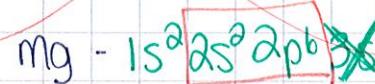
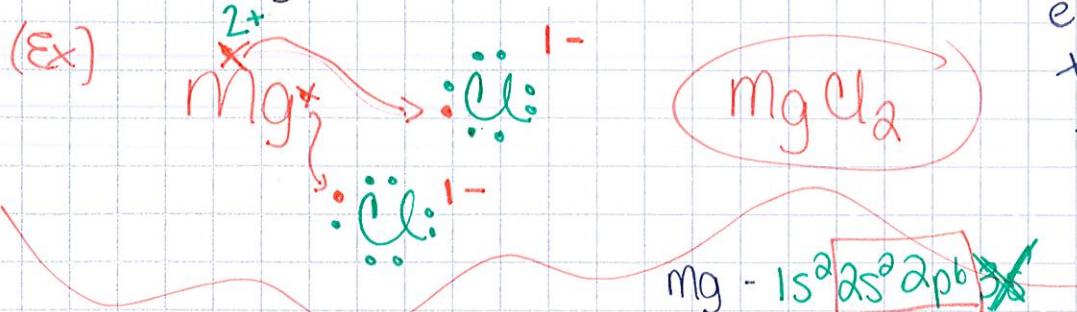


# review - e<sup>-</sup> dot diagrams

from Thursday



↳ showing formation of ionic bond



e<sup>-</sup> are transferred from metals to nonmetals to form the ionic bond.

## Properties of Ionic Compounds

- forming an ionic bond releases a lot of energy as heat (**exothermic**) -
- crystalline structures

Ex) sodium chloride



- it takes a lot of energy to melt/boil an ionic compound
  - high melting & boiling pts
- (Ex) NaCl m.p.  $800^{\circ}\text{C}$  ( $1472^{\circ}\text{F}$ )  
b.p.  $1413^{\circ}\text{C}$  ( $2575.4^{\circ}\text{F}$ )
- conduct electricity when dissolved in solution or molten

