

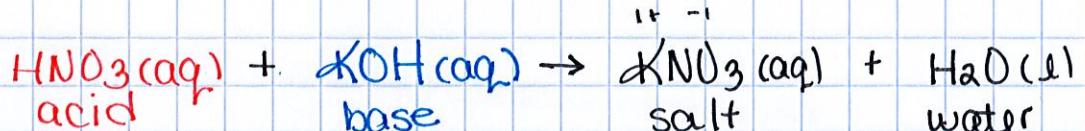
Properties of Acids & Bases

Acids

- taste sour
- electrolytes - conduct electricity
- turns blue litmus paper red
- react w/ metals to form Hz (g)
- react w/ bases to form a salt & water

Bases

- taste bitter
- electrolytes
- turn red litmus paper blue
- react w/ acids to form a salt & water

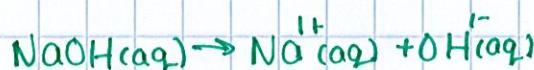
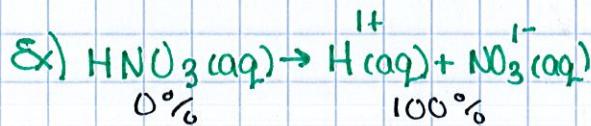


Strength of Acids & Bases

↳ refers to the extent to which an acid or base ionizes (breaks apart into ions) in solution, NOT how dangerous it is.

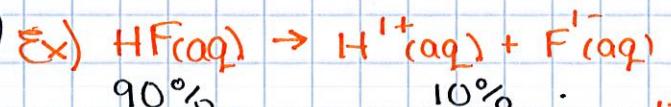
Strong Acids/Bases

- break apart (ionize) 100%



Weak Acids/Bases

- break apart (ionize) 10% or less



7 Strong Acids

- hydrochloric acid (HCl)
- hydrobromic acid (HBr)
- hydroiodic acid (HI)
- nitric acid (HNO_3)
- chloric acid (HClO_3)
- perchloric acid (HClO_4)
- sulfuric acid (H_2SO_4)

8 Strong Bases

- lithium hydroxide (LiOH)
- sodium hydroxide (NaOH)
- potassium hydroxide (KOH)
- rubidium hydroxide (RbOH)
- cesium hydroxide (CsOH)
- calcium hydroxide (Ca(OH)_2)
- strontium hydroxide (Sr(OH)_2)
- barium hydroxide (Ba(OH)_2)