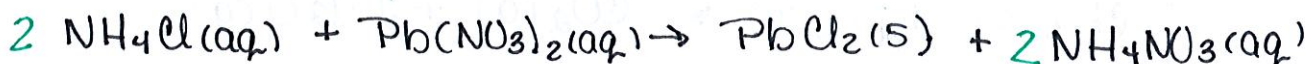


Key

Chemical Reactions Review

Directions: Write the skeleton equations, balance them, and determine the type of reactions for each problem.

1. Aqueous solutions of ammonium chloride and lead (II) nitrate react to form a precipitate of lead (II) chloride and a solutions of ammonium nitrate.



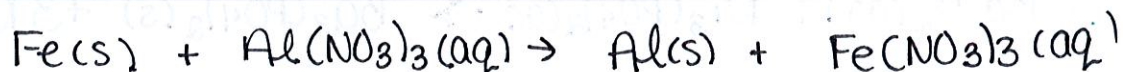
double-replacement

2. Water decomposes into hydrogen and oxygen gases when electricity is applied.



decomposition

3. Iron metal reacts with a solution aluminum nitrate to produce aluminum metal and a solution of iron (III) nitrate.



already balanced! single-replacement

4. Solid potassium nitrate breaks down in to solid potassium nitrite and oxygen gas.



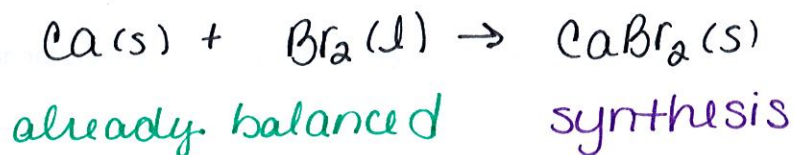
decomposition

5. Fluorine gas reacts with a solutions of potassium chloride to produce chlorine gas and a solution of potassium fluoride.

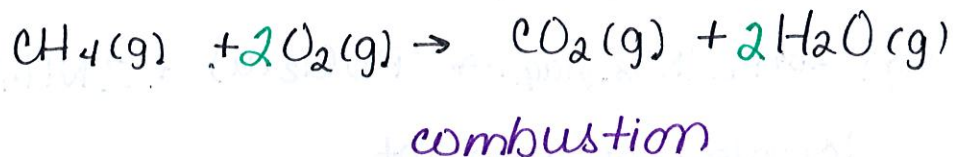


single-replacement

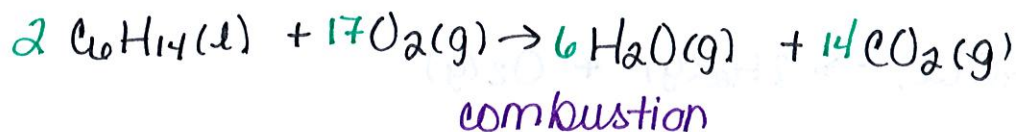
6. Calcium metal reacts with liquid bromine to create solid calcium bromide.



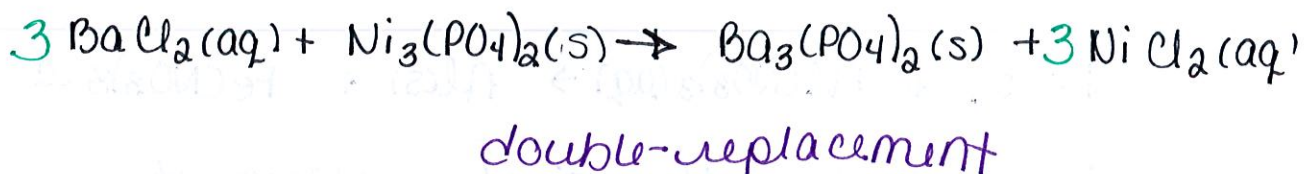
7. Carbon tetrahydride gas reacts with oxygen gas to produce carbon dioxide gas and water vapor.



8. Hexane (C_6H_{14}), a liquid, reacts with oxygen gas to produce water vapor and carbon dioxide gas.



9. A solutions of barium chloride reacts with solid nickel (II) phosphate to produce a precipitate of barium phosphate and a solution of green nickel (II) chloride.



10. Solutions of mercury (I) nitrate and gallium sulfate react to form a solutions of gallium nitrate and a precipitate of mercury (I) sulfate.

