

Periodic Table - arranged in order of increasing atomic #

- 118 elements - 98 occur naturally
- Some elements' symbols do NOT match their name b/c they match the element's Periodic Table original (Latin) name.

Name: _____

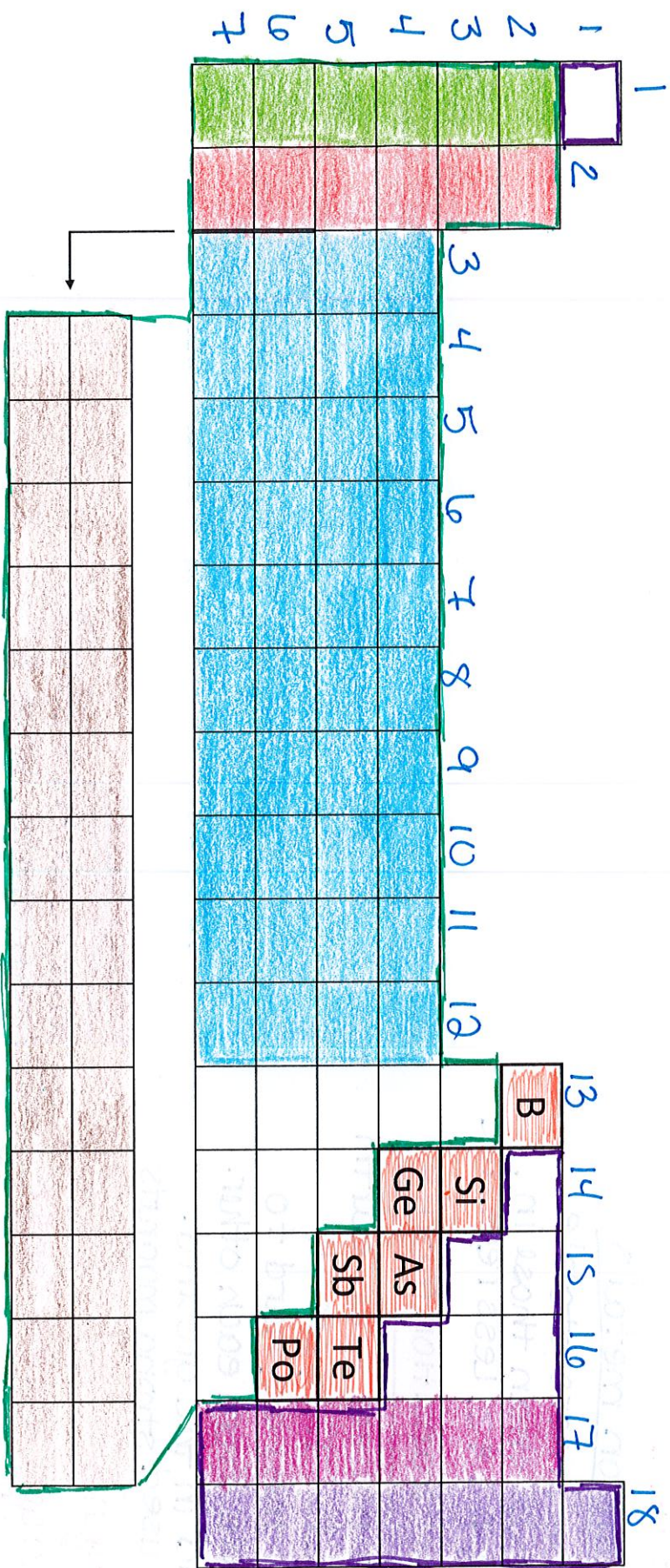
Legend

metals
nonmetals
metalloids
alkali metals
alkaline earth metals
transition metals

inner transition metals
halogens
noble gases

group: vertical column - all elements have same # valence e⁻

period: horizontal row



metals - lustrous, conductors of heat & electricity, malleable, ductile (drawn into wire)

nonmetals - either dull, brittle solids, liquid (br), or gases at room temperature. Nonconductors of electricity. metalloids - have characteristics of both metals & nonmetals

Alkali metals

most reactive metals, all have 1 valence e^- , called alkali metals b/c they form alkaline (basic) solutions in water.

Alkaline earth metals

almost as reactive as alkali metals, have 2 valence e^- , called so b/c they also produce alkaline solution & are found as minerals in the ground.

Transition metals

harder, more stable metals than those in groups 1 & 2. less reactive

Inner transition metals

also called rare earth metals. They aren't rare, they are hard to isolate from each other. Found in the ground. Also used strong magnets

Halogens

halogen means "salt former". most reactive nonmetals, have 7 valence e^- . commonly react w/ metals to produce salts (halides)

Noble gases

inert (non-reactive) nonmetals. Have 8 valence e^- . called noble because they do not interact w/ other elements like the nobility didn't interact w/ commoners.