

Key

Unit 7 Kinetics and Equilibrium Test Review

Kinetics

1. Any change that happens over a period of time can be expressed as a rate.
2. What is reaction rate? speed at which reactants react to form products.
3. Factors Affecting Reaction Rate
 - a. True or False: One way to observe the rate of a reaction is to observe the changes in [products] over time.
 - b. True or False: The rate of any reaction is a constant that does not change when reaction conditions (temperature, concentration, etc.)
 - c. Generally, an increase in temperature will increase the reaction rate.
 - d. True or False: Storing milk in the fridge stops the reactions that would cause the milk to spoil. It slows the reactions, but doesn't stop them.
 - e. How does an increase in surface area affect the exposure of reactants to one another? How does that affect the reaction rate? more reactant available to react. It increases collisions, which increase rates
 - f. True or False: Increasing the concentration of reactants will generally slow down the reaction.
 - g. True or False: A piece of material dipped in a concentrated dye solution will change color more quickly than in a dilute solution.
 - h. Why does an increase in pressure speed up the rate of a reaction? Particles collide more often
 - i. What is a catalyst? substance that speeds up a reaction by lowering the activation energy needed.
 - j. True or False: Because a catalyst is quickly consumed in a reaction, it must be added to the reaction over and over again to keep the reaction going? Catalysts are NOT consumed in a reaction
 - k. In your own words, explain why lowering the temperature slows down a reaction. Particles have less energy & move slower, collide less often
 - l. In your own words, explain why increasing the concentration of reactants will speed up a reaction? more particles collide more often

- m. Complete the following table by writing either increase or decrease for the rate of the reaction.

Scenario	Increase or Decrease
Adding heat	Increase
Removing heat	Decrease
Adding a catalyst	Increase
Diluting a solution (lowers concentration)	Decrease
Removing an enzyme (catalyst)	Decrease
lowering the temperature	Decrease
decreasing the surface area	Decrease
increasing the concentration of a solution	Increase
breaking a reactant down into smaller pieces	Increase

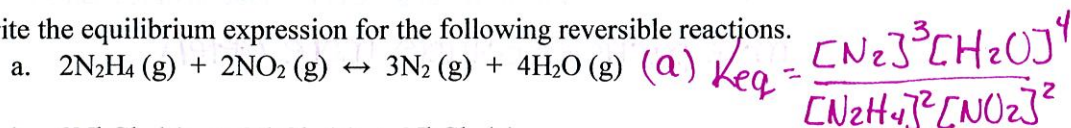
(increases surface area)

- n. Complete the following table by indicating which factor would have the greatest impact on the rate of the reaction. Choose from concentration, temperature, surface area, or catalyst.

Scenario	Factor that has the greatest impact
Blowing air on a campfire to help get it going.	concentration
Raw carrots are cut into thin slices for cooking.	surface area
Protein is broken down in the stomach by the enzyme pepsin.	catalyst
A woolly mammoth is found, perfectly preserved, near the Arctic Circle.	temperature
More bubbles appear when a concentrated solution of hydrochloric acid is added to a magnesium strip than when a dilute solution of acid is added.	concentration
Exhaust from a car engine passes through a catalytic converter changing most of the poisonous carbon monoxide to carbon dioxide.	catalyst
A dust explosion occurs in a saw mill.	surface area

Equilibrium

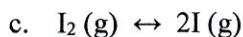
4. Write the equilibrium expression for the following reversible reactions.



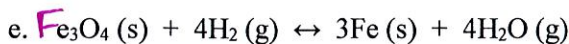
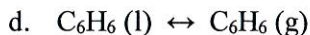
(b) $K_{eq} = \frac{[\text{NbCl}_3] [\text{NbCl}_5]}{[\text{NbCl}_4]^2}$

$$(c) K_{eq} = \frac{[I]^2}{[I_2]}$$

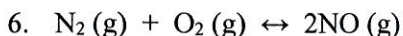
$$(d) K_{eq} = [C_6H_6]$$



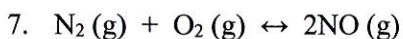
$$(e) K_{eq} = \frac{[H_2O]^4}{[H_2]^4}$$



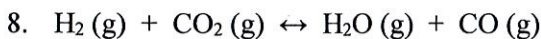
Calculate K_{eq} when $[PCl_5] = 0.0189 M$, $[PCl_3] = 0.0222 M$, and $[Cl_2] = 0.1044 M$. Is the forward or reverse reaction favored?



Calculate K_{eq} when $[N_2] = 1.01 M$, $[O_2] = 1.10 M$, and $[NO] = 0.999 M$.



$K_{eq} = 6.2 \times 10^{-4}$ and $[N_2] = 0.05200 M$ and $[O_2] = 0.00120 M$. Calculate $[NO]$.



$K_{eq} = 1.60$. Calculate the $[H_2]$ when $[CO_2] = .320 M$, $[H_2O] = .240 M$, and $[CO]$.

9. Using LeChatelier's Principle to fill in the chart.

Stress	$[N_2]$	$[H_2]$	$[NH_3]$	Equilibrium Shift
$\uparrow [N_2]$				
$\uparrow [H_2]$				
$\uparrow [NH_3]$				
\uparrow temperature				
$\downarrow [N_2]$				
\downarrow volume				
$\downarrow [NH_3]$				
$\downarrow [H_2]$				
add a catalyst				
\uparrow volume				
\downarrow temperature				

$$(5) K_{eq} = \frac{[PCl_3][Cl_2]}{[PCl_5]}$$

$$K_{eq} = \frac{[0.0222 M][0.1044 M]}{[0.0189 M]}$$

$$K_{eq} = .123 M$$

reverse is favored

$$(6) K_{eq} = \frac{[NO]^2}{[N_2][O_2]} = \frac{[0.999 M]^2}{[1.01 M][1.10 M]} = .898$$

reverse is favored

$$(7) K_{eq} = \frac{[NO]^2}{[N_2][O_2]}$$

$$6.2 \times 10^{-4} = \frac{[NO]^2}{[0.05200 M][0.00120 M]}$$

$$6.2 \times 10^{-4} = \frac{[NO]^2}{6.24 \times 10^{-5} M^2}$$

$$3.87 \times 10^{-8} = [NO]^2$$

$$\sqrt{3.87 \times 10^{-8}} = [NO]$$

reverse is favored

$$(8) K_{eq} = \frac{[H_2O][CO]}{[H_2][CO_2]}$$

$$1.60 = \frac{[.240 M][.240 M]}{[H_2][.320]}$$

$$1.60 = \frac{[.0576]}{[H_2][.320]}$$

$$1.60 \cdot [H_2] = \frac{[.0576]}{[.320]}$$

$$[H_2] = \frac{[.0576]}{[.320][1.60]} = .113 M$$

forward is favored