

# Le Châtelier's Principle

## Chem Worksheet 18-6

Name \_\_\_\_\_

A stress that is applied to a reaction that is at equilibrium conditions will shift the equilibrium position in a direction that tends to reduce this stress. This concept was first described by Le Châtelier. A reaction can be 'stressed' by changing the concentration of a reactant or product, changing the volume, and changing the temperature. Each stress tends to either favor the forward or reverse reaction until a new equilibrium position is established. If the forward reaction is increased we say equilibrium shifts to the right, and if the reverse reaction is increased equilibrium shifts to the left.

When the **concentration** of a gaseous or aqueous reactant or product is increased the equilibrium reaction shifts in the direction that decreases the concentration of that substance. If more product is added to a system at equilibrium, the reverse reaction increases in order to use the extra product, shifting equilibrium to the left. When the **volume** is reduced the equilibrium reaction shifts toward the side that contains the fewest gas particles. An increase in volume shifts to the side with the most gas particles. An increase in **temperature** will favor a reaction that is endothermic. A decrease in temperature will favor the reaction that is exothermic.

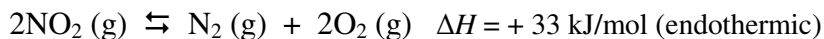
Concentration Changed	
$2A(g) + 3B(g) \rightleftharpoons 2C(g) + D(g)$	
Action	Effect
Increase [A]	Shift to the right
Increase [C]	Shift to the left
Decrease [B]	Shift to the left
Decrease [C]	Shift to the right

Temp. Changed for Endothermic Rxn.	
$A(g) + B(g) + \text{heat} \rightleftharpoons C(g) + D(g)$	
Action	Effect
Increase Temp.	Shifts to right (endothermic rxn.)
Decrease Temp.	Shifts to left (exothermic rxn.)

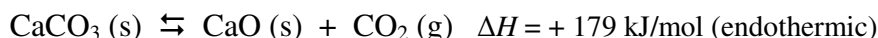
Volume Changed	
$2A(g) + 3B(g) \rightleftharpoons 2C(g) + D(g)$	
Action	Effect
Volume decreased	Shifts to right (side with fewest gases)
Volume increased	Shifts to left (side with most gases)

Temp. Changed for Exothermic Rxn.	
$A(g) + B(g) \rightleftharpoons C(g) + D(g) + \text{heat}$	
Action	Effect
Increase Temp.	Shifts to left (endothermic rxn.)
Decrease Temp.	Shifts to right (exothermic rxn.)

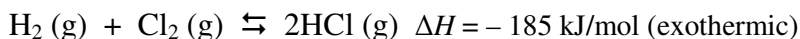
**Predict the direction equilibrium will shift when the following stresses occur. Explain your prediction. Assume each reaction occurs in a sealed container and has reached equilibrium.**



- NO<sub>2</sub> is added to the system.
- N<sub>2</sub> is added to the system.
- O<sub>2</sub> is removed from the system.
- The temperature of the container is increased.
- The volume of the container is increased.
- N<sub>2</sub> is added and NO<sub>2</sub> is removed.



- CO<sub>2</sub> is added to the system.
- The volume of the container is decreased.
- CaO is removed from the system.
- The temperature of the container is decreased.
- The volume of the container is increased.
- CaCO<sub>3</sub> is added to the system.



- H<sub>2</sub> is removed from the system.
- HCl is removed from the system.
- The volume of the container is increased.
- The temperature of the container is increased.
- The concentration of Cl<sub>2</sub> is decreased.
- The volume of the container is decreased.