

Mixed NOMENCLATURE Practice

Molecular Compounds, Ionic Compounds, & Acids

NAME THE FOLLOWING COMPOUNDS:

1. BaSO_3
2. $(\text{NH}_4)_3\text{PO}_4$
3. PBr_5
4. MgSO_4
5. CaO
6. H_3PO_4
7. $\text{Na}_2\text{Cr}_2\text{O}_7$
8. MgO
9. SO_3
10. $\text{Cu}(\text{NO}_3)_2$
11. HI
12. N_2O
13. MnO
14. $\text{C}_2\text{H}_5\text{OH}$
15. C_9H_{20}
16. C_5H_{12}
17. AgNO_3
18. As_2O_5
19. Fe_2O_3
20. HClO
21. N_2O_3
22. HF
23. $\text{H}_2\text{C}_2\text{O}_4$
24. NaHCO_3
25. SiBr_4
26. CuCl_2
27. HNO_2
28. SnO_2
29. BaCrO_4
30. $\text{C}_4\text{H}_9\text{OH}$

WRITE FORMULAS FOR THE FOLLOWING COMPOUNDS:

31. hydrobromic acid
32. chromium(III) carbonate
33. magnesium sulfide
34. iodine trichloride
35. lithium hydride
36. ammonium hydroxide
37. calcium chloride
38. hydroselenic acid
39. iron(II) nitride
40. aluminum hydroxide
41. tin(II) fluoride
42. sulfur tetrachloride
43. mercury(II) iodide
44. hexane
45. propanol
46. diphosphorus pentoxide
47. sulfurous acid
48. lead(II) nitrate
49. dihydrogen monoxide
50. sodium oxalate
51. perchloric acid
52. chlorous acid
53. silicon dioxide
54. carbonic acid
55. sodium chlorate
56. methane
57. heptanol
58. xenon hexafluoride
59. nickel nitrate
60. potassium perchlorate