

Name:

Periodic Table

Legend

- transition metals**: Yellow
- halogens**: Blue
- noble gases**: Red
- alkali metals**: Orange
- alkaline earth metals**: Red-orange
- inner transition metals (rare earth metals)**: Green
- metalloids**: Purple
- nonmetals**: Pink
- metals**: Light blue

horizontal rows: vertical columns

group: Other metals

period: Other nonmetals

	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1																		
2																		
3																		
4																		
5																		
6																		
7																		

Lanthanides (Actinides)

metals - lustrous, malleable, ductile, conduct heat & electricity

nonmetals - dull, brittle, nonconductors (insulators)

metalloids - have properties of both metals & nonmetals

Alkali metals - called so b/c they form alkaline (basic) solutions in water

- most reactive metals
- do not occur free in nature
- have 1 valence e⁻
- soft metals

Alkaline Earth metals - very reactive metals

- not found free in nature
- have 2 valence e⁻

Transition Metals - harder, less reactive than groups 1 & 2

- Fe, Co, & Ni - produce magnetic fluids
- Cu, Ag & Au - coinage metals

Halogens - most reactive nonmetals

- name means "salt former"
- have 7 valence e⁻

Noble Gases - considered inert, nonreactive, nonmetals

- have 2 valence e⁻ (He) or 8 valence e⁻ (the rest)
- very stable

Inner Transition metals - (rare earth metals)

- considered rare b/c they are very hard to separate