$q = mC\Delta T$ $\Delta T = T_f - T_i$

#1. A. 15.75 g piece of iron absorbs 1086.75 J of heat and its temperature rises from 25.00°C to 175.0°C. Calculate the specific heat capacity of iron.

$q = mC\Delta T$ $\Delta T = T_f - T_i$

#2. How many J of heat are needed to raise the temperature of 10.0 g of aluminum from 22.2°C to 55.5°C?

$\Delta T = T_f - T_i$

piece of iron absorbs
J of heat and its
rature rises from 25.00°C
5.0°C. Calculate the
ific heat capacity of iron.

$q = mC\Delta T$ $\Delta T = T_f - T_i$

#3 100.0g of 4.00°C water is heated until its temperature is 37.00°C. Calculate the amount of heat absorbed.

$q = mC\Delta T$ $\Delta T = T_f - T_i$

#4 Calculate the specific heat capacity of a piece of wood if 1500.0g releases 67500J of heat as it cools from 57.1°C to 32.0°C.

$q = mC\Delta T$ $\Delta T = T_f - T_i$

#5. 55.00g of silicon releases heat as its temperature falls by 15.00°C. How much heat is released

$q = mC\Delta T$ $\Delta T = T_f - T_i$

#6 What mass of ice will change its temperature by 3.000°C when 525J of heat is added?

$\Delta T = T_f - T_i$

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5.0°C. Calculate the
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#7 A 0.303g piece of brass is heated and fashioned into a bracelet. What was the change in temperature when 46.300 J of heat is added to the ~~copper~~ brass

$q = mC\Delta T$ $\Delta T = T_f - T_i$

#8 25.0g of mercury is cooled from 155.0°C to 25.00°C and releases 455 J of heat. Calculate the specific heat capacity of mercury