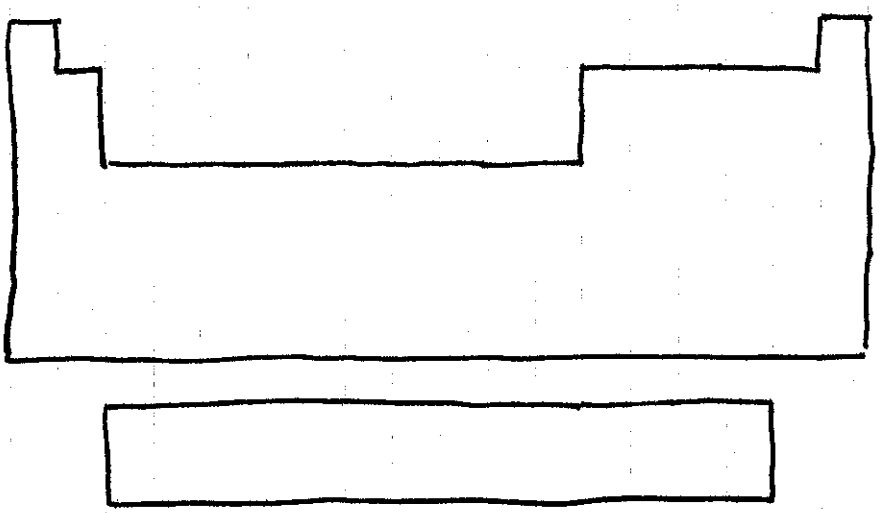
Unit 2 Review



1) Number the groups. Define a group.

2) Number the periods. Define a period.

3) Outline the following (and label): metals, metalloids, & nonmetals

4) Label these groups:
   - noble gases
   - alkali metals
   - alkaline earth metals
   - transition metals
   - inner transition metals
   - halogens

5) Who is credited for developing the 1st Periodic Table? How did he arrange his Periodic Table?

6) How is the modern Periodic Table arranged?

7) Define:
   - atomic radius
   - ionization energy
   - electronegativity

8) Explain the following trends:

	atomic radius	ionization energy	electronegativity
down a group			
across a period			

9) Use your knowledge of trends to answer these questions.

- which element has a larger atomic radius: Cr or Cu?
- which element has a larger ionization energy: Cr or W?
- order these elements from smallest to largest electronegativity.

Si Fl Pb } Ag Mo Sb

- order these elements in decreasing ionization energy.

Si Fl Pb } Ag Mo Sb

- order these elements in increasing atomic radius

Si Fl Pb } Ag Mo Sb

10) Study your answers to the "Hunting the Elements" Documentary.