Notes: % Composition

% by mass of each element in a compound

$$\% \text{ element} = \frac{\text{mass element}}{\text{molar mass}} \times 100$$

&) Find the % of each element in OF_2

1 O (16.00g) = 16.00g

2 F (19.00g) = 38.00g

54.00g

$$\%O = \frac{16.00\,g}{54.00\,g} \times 100 = 29.63\%$$

$$\%F = \frac{38.00\,g}{54.00\,g} \times 100 = 70.37\%$$

Find the % of each element in $Ca(OH)_2$

%Ca = 54.09%

%O = 43.18%

%H = 2.73%

54.09
43.18
2.73