

Type III Covalent Nomenclature

2 nonmetals
sharing e⁻

Use Prefixes to show # atoms of each element.

1 mono-	6 hexa-
2 di-	7 hepta-
3 tri-	8 octa-
4 tetra-	9 nona-
5 penta-	10 deca-

Formulas

- Write the symbol for the 1st element, change the prefix to a subscript.
- Do same for 2nd element.
- **Newer, ever criss-cross!**

ex) dihydrogen monoxide



iodine heptabromide



sulfur trioxide

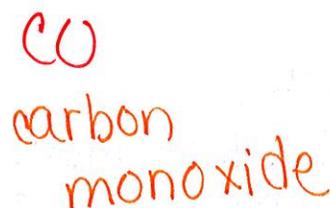
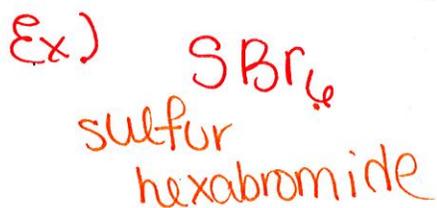


dichlorine heptasulfide

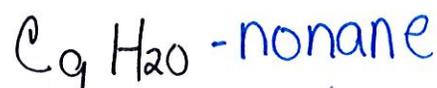
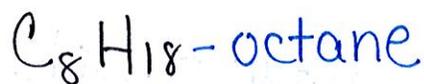
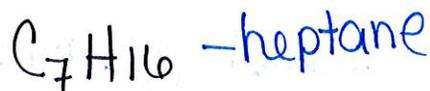
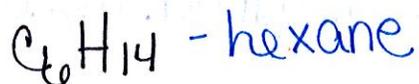
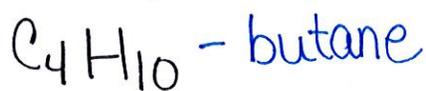
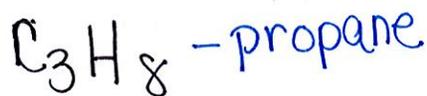
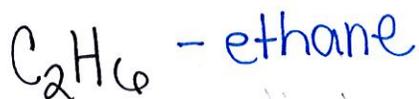


Names

- Write name 1st element, change subscript to a prefix
- Do same for 2nd element, change ending to -ide.



Simple Alkanes (Organic) Hydrocarbons



Acidic Nomenclature

All acids have H^{1+} as the first ion in the formula.

Binary Acids - Do NOT have polyatomic ions w/oxygen in them!

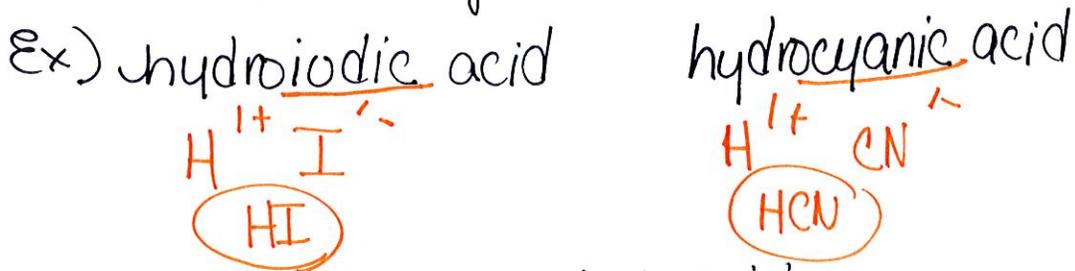
Names

- write the prefix HYDRO-
- write the anion's name, change ending to -ic acid.

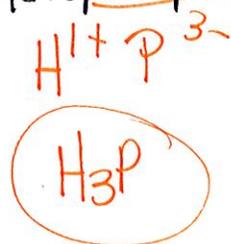


Formulas

- write H^{1+}
- write symbol & charge of the anion
- Criss-cross charges



hydrophosphoric acid.



Oxyacids - have polyatomic ions w/
oxygen in them!

Names - newer uses hydro-

look @ the name for the
polyatomic ion.

ends in ate
change ate to ic acid

ends in ite
change ite to ous acid



nitric
acid



chlorous
acid



oxalic
acid



phosphorous
acid

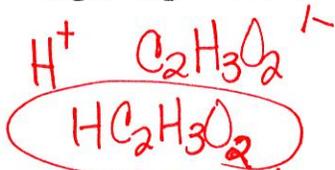
Formulas

• write H^{1+}

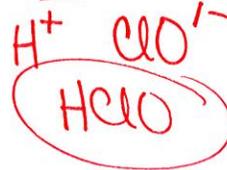
• If acid name ends in ic use poly. ion that
ends ate. If name of acid ends -ous, use
poly. ion that ends in ite

• Criss-cross

Ex) acetic acid



hypochlorous acid



sulfurous acid

