

Basic Atomic Structure Worksheet

1. The 3 particles of the atom are:

a. _____
b. _____
c. _____

Their respective charges are:

a. _____
b. _____
c. _____

2. The number of protons in one atom of an element determines the atom's _____, and the number of electrons determines the _____ of the element.
3. The atomic number tells you the number of _____ in one atom of an element. It also tells you the number of _____ in a neutral atom of that element. The atomic number gives the "identity" of an element as well as its location on the periodic table. No two different elements will have the _____ atomic number.
4. The _____ of an element is the average mass of an element's naturally occurring atom, or isotopes, taking into account the _____ of each isotope.
5. The _____ of an element is the total number of protons and neutrons in the _____ of the atom.
6. The mass number is used to calculate the number of _____ in one atom of an element. In order to calculate the number of neutrons you must subtract the _____ from the _____.
7. Give the symbol of and the number of protons in one atom of:

Lithium _____
Iron _____
Oxygen _____
Krypton _____

Bromine _____
Copper _____
Mercury _____
Helium _____

8. Give the symbol of and the number of electrons in a neutral atom of:

Uranium _____
Boron _____
Chlorine _____

Iodine _____
Xenon _____

9. Give the symbol of and the number of neutrons in one atom of:

(Mass numbers are ALWAYS whole numbers...show your calculations)

Barium _____
Carbon _____
Fluorine _____
Europium _____

Bismuth _____
Hydrogen _____
Magnesium _____
Mercury _____

10. Name the element which has the following numbers of particles:

- 26 electrons, 29 neutrons, 26 protons _____
- 53 protons, 74 neutrons _____
- 2 electrons (neutral atoms) _____
- 20 protons _____
- 82 electrons, 125 neutrons, 82 protons _____
- 0 neutrons _____

11. If you know ONLY the following information can you ALWAYS determine what the element is? (Yes/No)

- Number of protons _____
- Number of neutrons _____
- Number of electrons in a neutral atom _____
- Number of electrons _____

12. Fill in the missing items in the table below.

NAME	SYMBOL	Z	A	# PROTONS	# ELECTRONS	# NEUTRONS	ISOTOPIC SYMBOL
a.	Na						
b.		17			17		
c. Potassium							
d.	P						
e. Iron					26		
f.				53			
g. Silver							
h.		36					
i.	W						
j.		29					
k.				49			
l.				79	79		
m.		16			16		